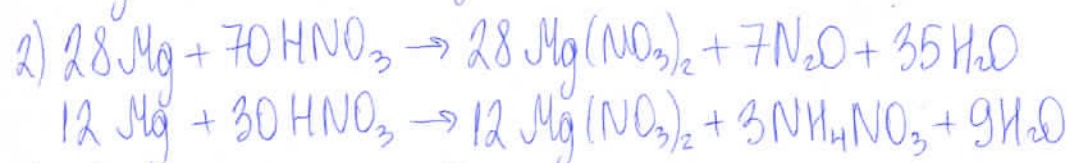
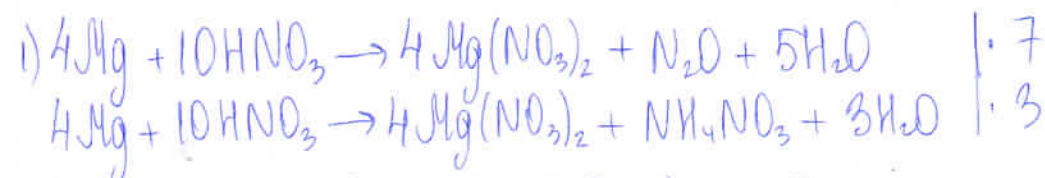


№1

55.



3) Всего 100 моль HNO_3 , израсход

$$n(HNO_3) = 0,17 \cdot 100 = 17 \text{ моль}$$

Ответ: 17 моль

№2

55.

Дано:

$$m(SO_3) =$$

$$m(H_2SO_4) = 10 \text{ г}$$

$$W(SO_3) = 30\%$$

$$\rho(HNO_3) = 1,413 \text{ г/мл}$$

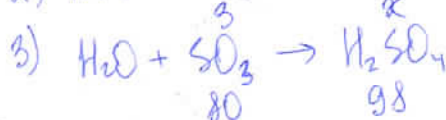
$$W(HNO_3) = 70\%$$

$$V(HNO_3) = ?$$

Решение:

$$1) m(SO_3) = \frac{10 \cdot 30\%}{100\%} = 3 \text{ г}$$

$$2) m(H_2SO_4) = 10 - 3 = 7 \text{ г}$$



$$x = \frac{3 \cdot 98}{80} = 3,675 \text{ г}$$

$$4) m(H_2SO_4) = 3,675 + 7 = 10,675 \text{ г}$$

$$5) m(HNO_3) = V(HNO_3) \cdot \rho(HNO_3) \cdot \frac{70\%}{100\%}$$

$$m(HNO_3) = V(HNO_3) \cdot 1,413 \cdot 0,7$$

$$6) \frac{W(HNO_3)}{W(H_2SO_4)} = \frac{1}{2} = \frac{m(HNO_3)}{m(H_2SO_4)}$$

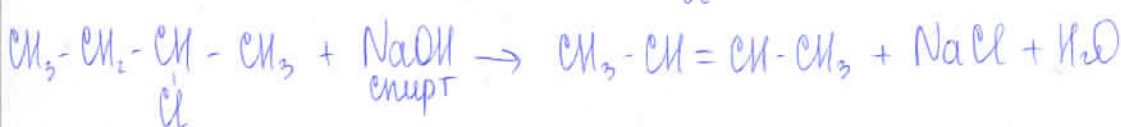
$$\frac{V(HNO_3) \cdot 1,413 \cdot 0,7}{10,675} = \frac{1}{2}$$

$$V(HNO_3) = \frac{10,675}{1,413 \cdot 0,7 \cdot 2} = 5,39 \approx 5,4 \text{ мл}$$

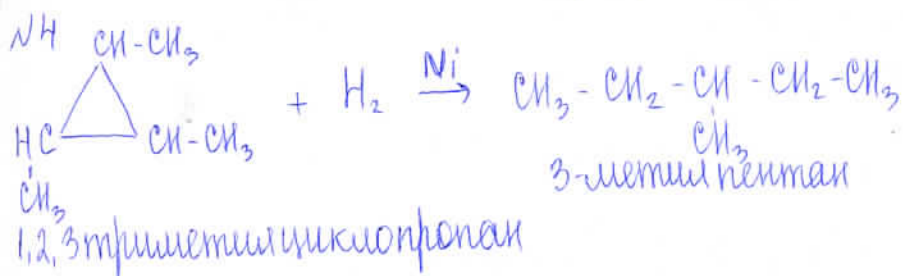
Ответ: 5,4 мл

№3

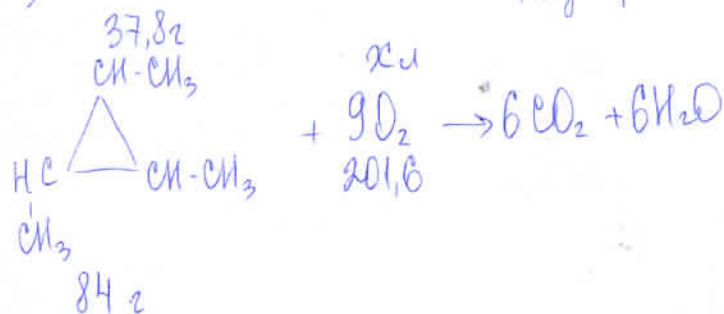
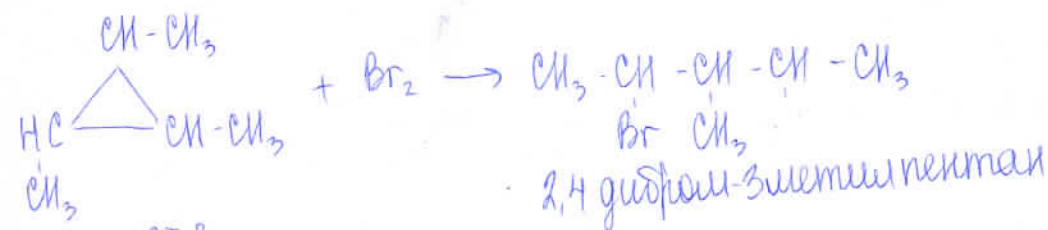
25.



Дугиен-1 → 2 курс | Дугиен → Дугиен-2.

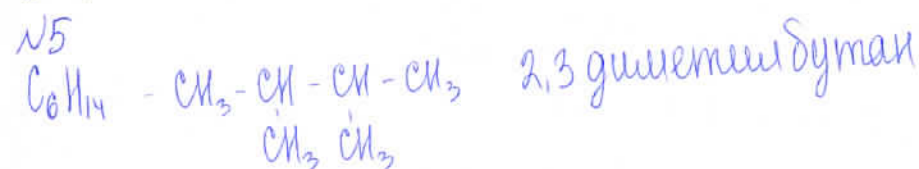


55.

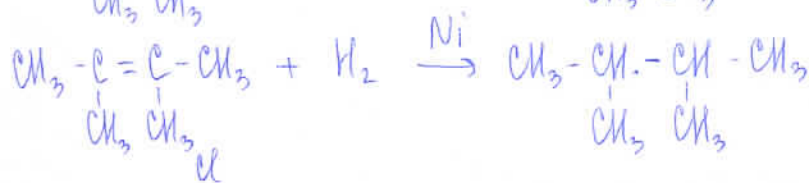
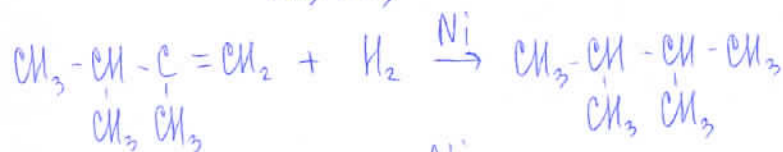


$$n = \frac{37,8}{84} = 90,72 \mu$$

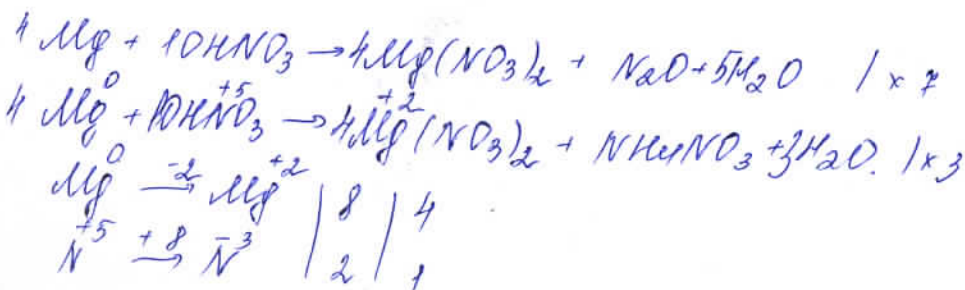
Ombem: 90,72 u



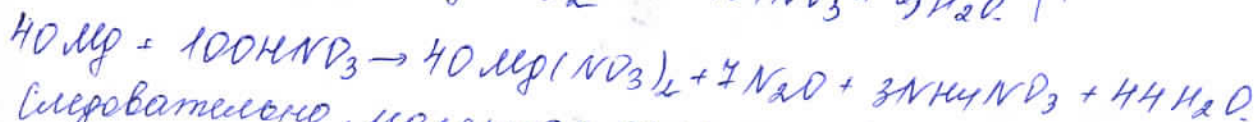
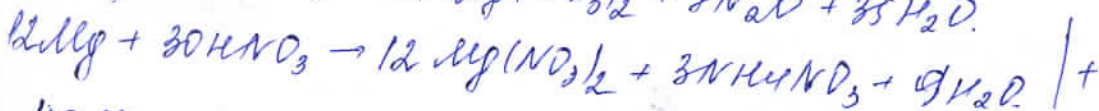
55.



1.



55.

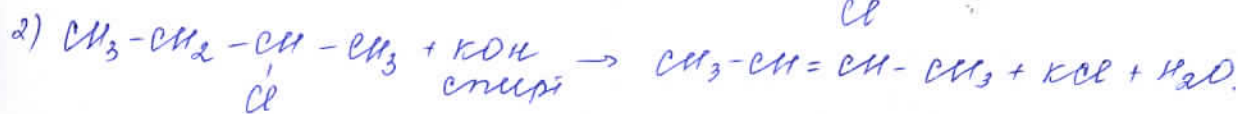


Следовательно, молярная доля азота (HNO_3) = 100 моль. \Rightarrow

$\Rightarrow 0,18 \cdot 100 = 18$ моль (HNO_3) первонач. восстановления (V)
 Ответ: 18 моль.

2.

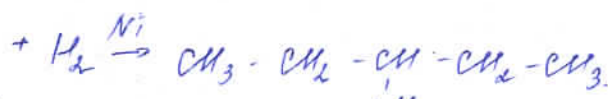
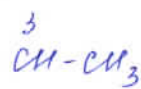
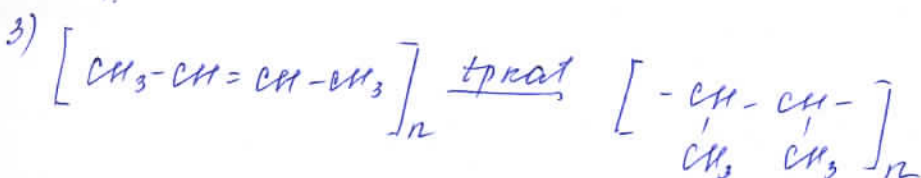
3.



55.

исходное в. во - бутен-1. - C_4H_8

конечное в. во - бутен-2. - C_4H_8
 $n=4$



(3-метилпентан)

55.

1, 2, 3-триметилциклопропан. - C_6H_{12}

Бензеновое Окисление 11.15^ч

170



55

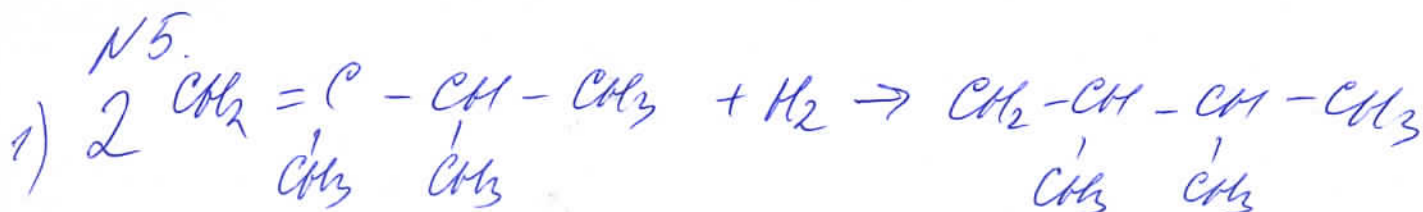
$$V(C_6H_{12}) = \frac{37,8}{84} = 0,45 \text{ моль}$$

$$V(CO_2) = 0,45 \cdot 9 \cdot 22,4 = 90,72 \text{ л}$$

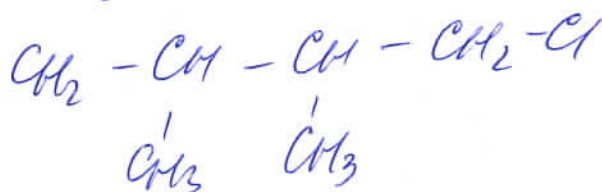
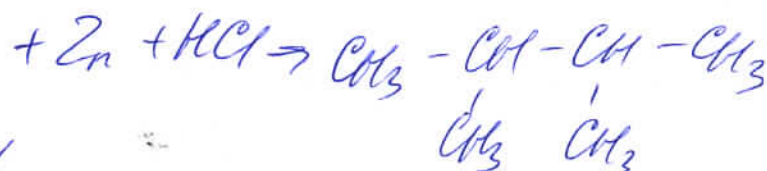
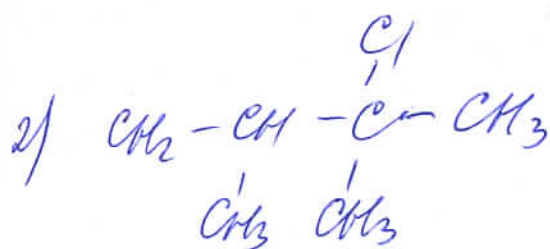
Ответ: 90,72 л.



25



58



N2. Дано:

$$m(\text{окисл}) = 102$$

$$m(SO_3) = 3,675$$

$$m(H_2SO_4) = 70,00\%$$

$$\rho = 1,413 \text{ г/мл}$$

$$V_{p-p} (HNO_3) = ?$$

P-e.



58.

$$V(SO_3) = \frac{3}{80} = 0,0375 \text{ моль}$$

$$m(H_2SO_4) = 98 \cdot 0,0375 = 3,675$$

$$m(H_2SO_4) = 7 + 3,675 = 10,675$$

$$\frac{W(H_2SO_4)}{W(HNO_3)} = 2$$

$$\frac{W(H_2SO_4)}{W(HNO_3)} = \frac{10,675}{1,413 \cdot V \cdot 0,7}$$